

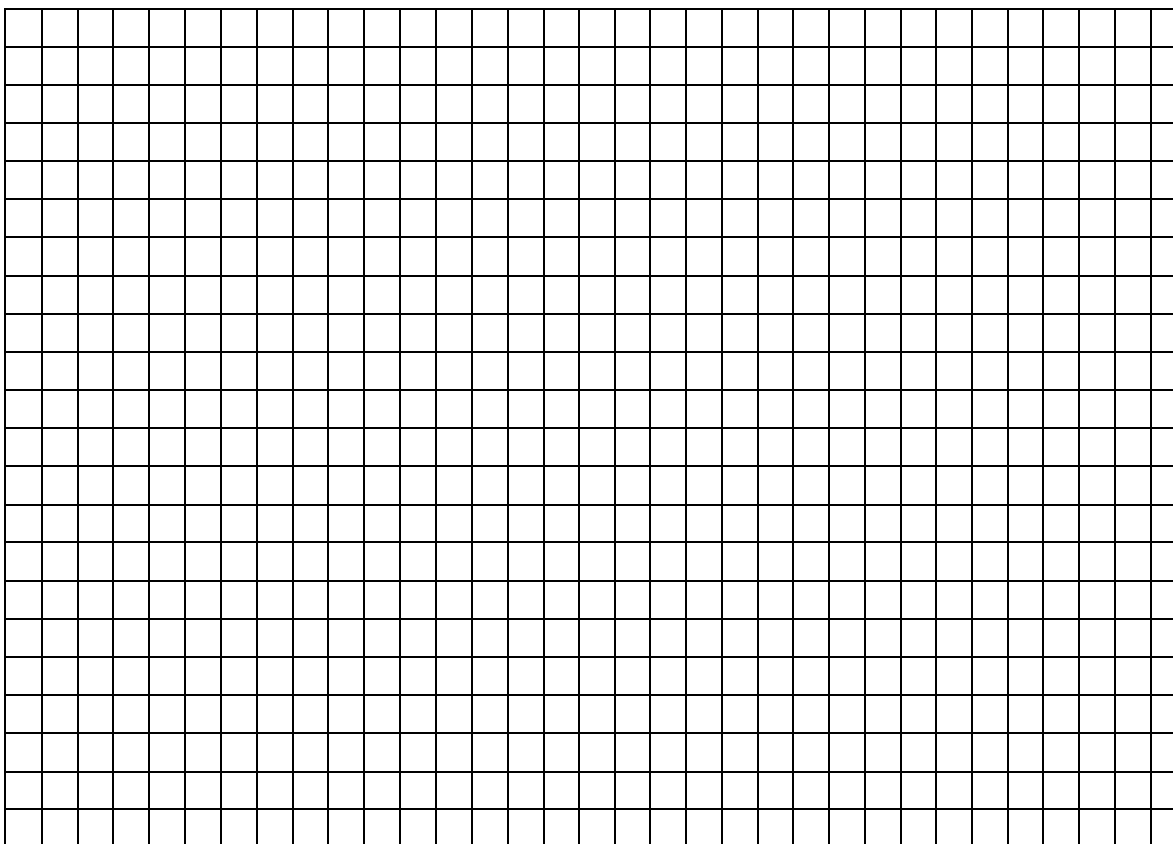
Gas Law I DCI

Charles Law

Using the simulation http://introchem.chem.okstate.edu/DCICLA/charles_law.swf
Collect temperature and volume data for five different points.

Temperature ($^{\circ}\text{C}$)	Temperature (K)	Volume (mL)

Plot the data.



Do you see any pattern in the data? Explain.

If you extrapolate the graph to $V = 0$, what would be the temperature in $^{\circ}\text{C}$ and K ?

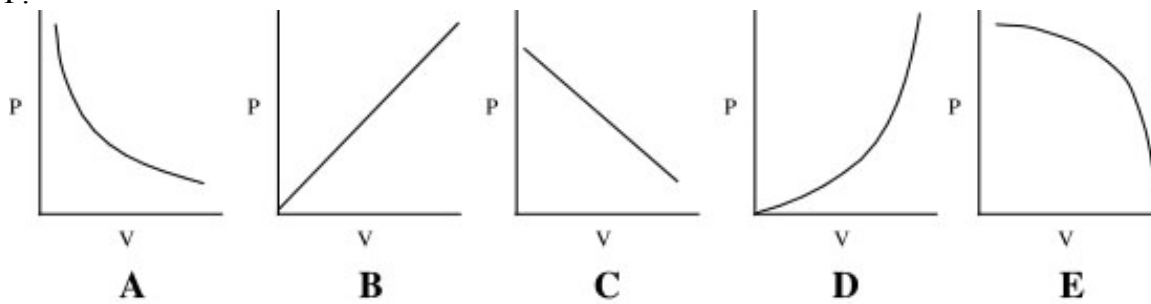
Write an algebraic equation representing the relationship between V and T for an ideal gas.

View the video at <http://genchem1.chem.okstate.edu/CCLIEMD09/Hemoviesmall.mov>

Describe what you see. Explain what is happening to the gas molecules in the demonstration.

In the space below use two particulate level drawings to represent the gas before cooling and then after cooling.

In the BCE you identified a graphical representation for the relationship between V and P .



Select from the graphs below a second way to represent the same inverse relationship between P and V .

